Experiment 2 Worksheet: Conservation of Copper

Purpose: (A statement should be made here regarding what your experiment is designed to accomplish	'n.)
Procedure: (Describe here how you went about completing your experiment. Use enough detail that someone could repeat it, if necessary. Also reference any books or manual you may have used to assist you.)	
Data:	
Before Reaction:	
Mass of 250 mL side-arm flask	
Mass of 5 or 6 pieces of solid zinc metal on Kimwipe	
Mass of 1 Kimwipe	
Mass of 250 mL side-arm flask with CuSO ₄ acidic solution	
Mass of CuSO ₄ acidic solution alone	
Mass of Zn Metal alone	
After Reaction:	
Mass of metal pieces on Kimwipe	
Mass of 250 mL side-arm flask with CuSO ₄ acidic solution	
Mass of solution alone Mass of metal pieces alone	
Calculated mass of H ₂ gas alone	
Total mass before reaction Total mass after reaction	_
Mass % recovered after reaction (Total Mass After Reaction *100%) Total Mass Before Reaction	_

Calculations: (Show all calculations used to produce the above data here.)
Conclusions: (Summarize your results here and discuss possible errors that may have occurred during the experiment. Does the mass you recovered support the law of conservation of mass? If it does, explain
the expertment. Does the mass you recovered support the law of conservation of mass? If it does, explain how it does. If it doesn't explain, why it does not. How do the possible errors in your experiment affect you decision as to whether or not your experiment supports the law?)