CHM 3120L

INTRODUCTION TO ANALYTICAL CHEMISTRY

LABORATORY REPORT

**EXPERIMENT: DETERMINATION OF ASCORBIC ACID BY REDOX TITRATION**

Name: Click here to enter text.

Section: Click here to enter text.

Date Experiment Completed: Click here to enter a date.

1. Complete the following table:

|  |  |
| --- | --- |
|  | **Preparation of the potassium iodate solution** |
|  | F.W., g/mole | 214.0008 |
|  | mass, g |  |
|  | volume, mL | 500.00 |
|  | concentration, mole/L |  |
|  | volume titrated, mL | 25.00 |
|  | moles titrated |  |

1. Complete the following table. (You also may create a similar table in Excel, complete your calculations there, and then paste it below. Feel free to add columns to the table to help you with your calculations).

|  |
| --- |
| **Calibration of the sodium thiosulfate solution** |
| Volume of Na2S2O3 solution used for titration, mL  | concentration, mole/L |
| Trial 1 | Trial 2 | Trial 3 | average |
|   |   |   |   |  |

1. Complete the following table. (You also may create a similar table in Excel, complete your calculations there, and then paste it below. Feel free to add columns to the table to help you with your calculations).

|  |
| --- |
| **Determination of the Ascorbic Acid** |
| Trial | mass of the sample, g | volume of Sodium Thiosulfate solution, mL | Weight % of the Asc. Acid in the tablet |
| 1 |   |   |   |
| 2 |   |   |   |
| 3 |   |   |   |

4. Report you final result:

Weight Percent: Click here to enter text.

Uncertainty: Click here to enter text.